

Chemistry
Form 3
Scheme of Work

TERM II

Week	Topic	Syllabus Objectives	Activities/ Labs
1-2	<u>Revision of exam papers</u>	To clear up misconceptions and other problems.	Students can view papers of other students & answers done on the board.
3-7	<u>Acids, bases & salts (continued)</u>	<p><u>Acids, bases & salts</u></p> <p>1: Describe the pH scale with aid of a diagram</p> <p>2: <u>Acids</u></p> <ul style="list-style-type: none"> • a): Describe the properties of acids • b): Determine the types of acids: <ul style="list-style-type: none"> • i) strong acids • ii) weak acids • c): Distinguish the position of strong/ weak acids on the pH table • d): List examples of strong/ weak acids • e): Identify the general reactions (rxns) of acids with: i) metals ii) bases or alkalis iii) carbonates <p>3: <u>Alkalis & bases</u></p> <ul style="list-style-type: none"> • Define alkali and Base • Describe the properties of alkalis • Determine the types of alkalis: 	<p>Reactions will be demonstrated in the lab and practice exercises done in class to ensure students can write balanced chemical equations given only parts of a word equation. (see Chemistry for You pg 155)</p> <p>ii) Describe the test for hydrogen and carbon dioxide when performing these tests</p> <p>iii) Perform rxns in the lab</p> <p>Course Work</p>

		<ul style="list-style-type: none"> • i) strong alkalis ii) weak alkalis • Distinguish the position of strong/ weak alkalis on the pH table • Main rxns. of alkalis with: i) metal ions ii) ammonium salts <p>4: <u>Salts</u></p> <ul style="list-style-type: none"> • Definition of a salt • Identify the types of salts eg. Chlorides, sulphates etc. from the strong acids • Introduce the solubility table and its use in writing state symbols in equations • Define a precipitate 	
8-10	<u>The Reactivity series</u>	<p>1: List the reactivity Series (R.S)</p> <p>2: Describe reactivity in descending order</p> <p>3: <u>Write chemical rxns for:</u></p> <p>i) water & K, Na, Ca</p> <p>ii) steam & Mg-Fe</p> <p>iii) acid & Mg-Fe</p> <p>4: Determine displacement rxns with a more reactive metal and a i) metal in solution ii) metal oxide and hence write balanced equations for such rxns (practice equations with state symbols)</p>	<p>i) Show the difference in the reactivity of some metals from the RS</p> <p>ii) Show displacement rxns eg. Mg & CuSO₄ etc.</p> <p>Course Work/ Science Fair Mark</p>